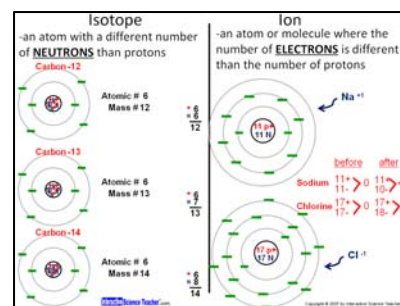


Isotopes And Ions

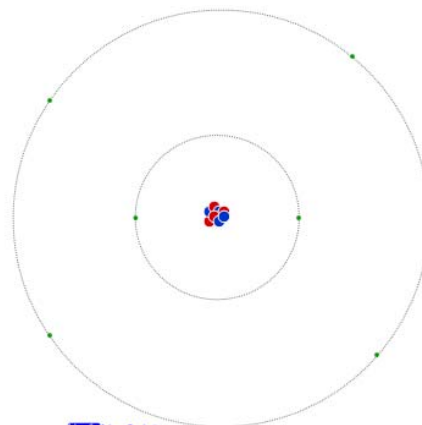
Confused about isotopes and ions? Let this lesson help straighten you out. 3 carbon isotopes are presented.

Then on the other side of the paper we'll watch chlorine steal an outer electron from sodium and create an ion.



1. Start the day with a little review. Go over the atom- protons and neutrons in the nucleus, with electrons zipping around it.

This entire lesson is available as a PowerPoint- see last page.



2. If you're using the Student Handout (see last page), have students copy the definition as shown. Otherwise, on a blank sheet of paper have them do so.

If you're using the PowerPoint, let students know their job is to copy everything they see on the screen.

<p>Isotope -an atom with a different number of NEUTRONS than protons</p>	
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Isotope

-an atom with a different number of **NEUTRONS** than protons

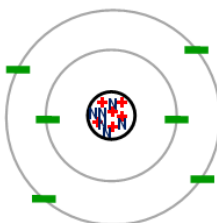
3. Below the Isotope definition, draw what will be the nucleus for our first isotope- carbon-12. Draw 6 little +’s and 6 little N’s together.



Isotope

-an atom with a different number of **NEUTRONS** than protons

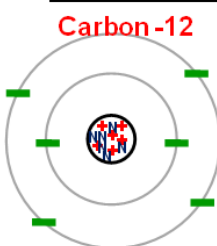
4. Add the electrons. We’ll use the Bohr model today to keep things a little more clear. If you’ve already covered the model of the atom, remind students that this has changed into the electron cloud model.



Isotope

-an atom with a different number of **NEUTRONS** than protons

5. With the atom now complete, we’ll count the number of protons (that’s our atomic number), number of neutrons (add to protons for mass number). The mass number 12 gets added to the end of the name “Carbon”.

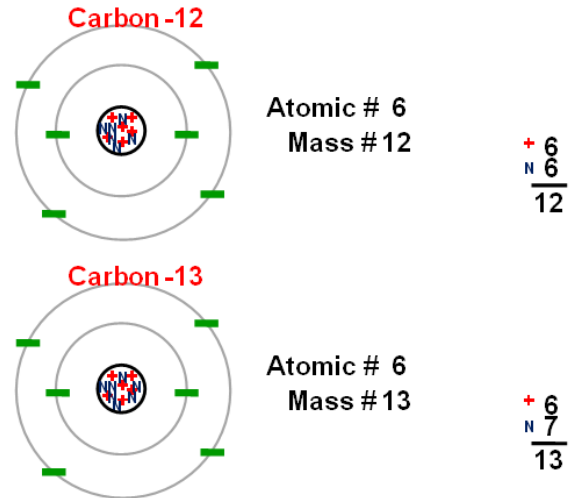


Carbon-12

Atomic # 6
Mass # 12

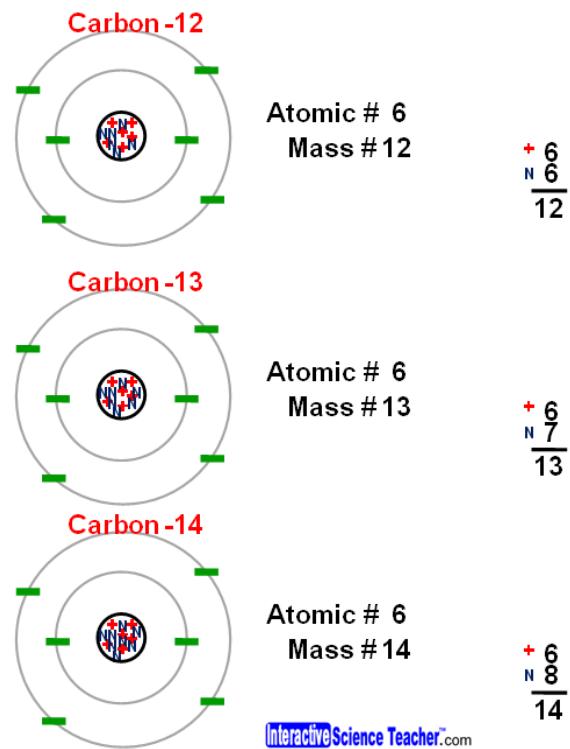


6. Let's make another atom that almost the same, except it has one more neutron in the nucleus. When you add up the particles in the nucleus, your new mass number is 13.



7. The last one (Carbon-14) has another extra neutron (8 total). That makes the mass number 14.

That's isotopes in a nutshell.



Isotope

-an atom with a different number of **NEUTRONS** than protons

8. Now onto ions, which we'll do on the other side of the paper. Start with a simple definition.

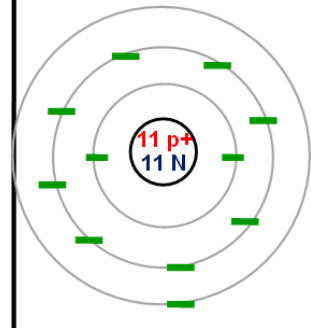
<p>type</p> <p>different number of protons</p>	<p>Ion</p> <p>-an atom or molecule where the number of ELECTRONS is different than the number of protons</p>
<p>atomic # 6</p> <p>mass # 12</p>	$\begin{array}{r} +6 \\ N \\ \hline 12 \end{array}$
<p>atomic # 6</p> <p>mass # 13</p>	$\begin{array}{r} +6 \\ N \\ \hline 13 \end{array}$
<p>atomic # 6</p> <p>mass # 14</p>	$\begin{array}{r} +6 \\ N \\ \hline 14 \end{array}$

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9. Draw an atom of sodium as shown. We won't draw every proton and neutron here because they aren't the point.

Ion

-an atom or molecule where the number of **ELECTRONS** is different than the number of protons

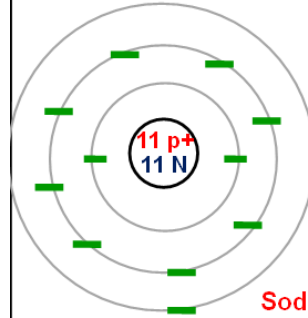


The diagram shows a central nucleus with 11 protons (p+) and 11 neutrons (N). It is surrounded by three concentric shells of electrons, represented by green dashes. The innermost shell has 2 electrons, the middle shell has 8 electrons, and the outermost shell has 1 electron.

10. Take a "before" inventory. Right now we have 11 protons and 11 electrons. The sum of those charges is zero.

Ion

-an atom or molecule where the number of **ELECTRONS** is different than the number of protons

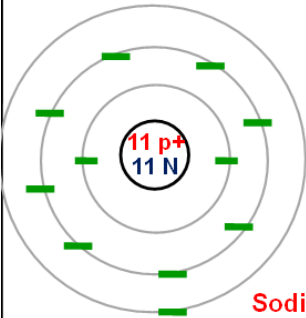


before after
 Sodium 11+ > 0
 11-

11. Below sodium draw an atom of chlorine. Take its inventory (17+ and 17- = 0).

Ion

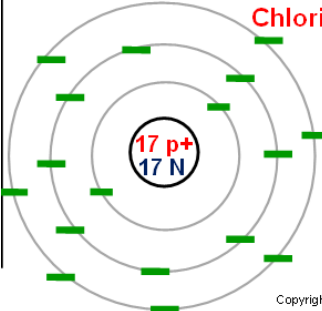
-an atom or molecule where the number of **ELECTRONS** is different than the number of protons



Sodium
11+
11- > 0

before after

Chlorine
17+
17- > 0




Chlorine
17+
17- > 0

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12. Notice that one lonely outer electron that sodium has, and the one electron that chlorine *could* use, and then move it.

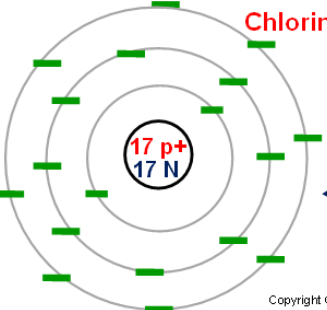
13. Do an “after” inventory with the new electron arrangement. Sodium will now be an overall +1 because it lost a negatively charged electron, and chlorine will be -1 because it gained the electron.



Sodium
11+
10- > +1

before after

Chlorine
17+
18- > -1



Cl⁻¹

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PowerPoint- lead your students through the quiz click-by-click

Isotope
-an atom with a different number of **NEUTRONS** than protons

Carbon-12
Atomic # 6
Mass # 12
 $\frac{+6}{+6} = 12$

Carbon-13
Atomic # 6
Mass # 13
 $\frac{+6}{+7} = 13$

Carbon-14
Atomic # 6
Mass # 14
 $\frac{+6}{+8} = 14$

Ion
-an atom or molecule where the number of **ELECTRONS** is different than the number of protons

Isotope
-an atom with a different number of **NEUTRONS** than protons

Carbon-12
Atomic # 6
Mass # 12
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Atomic # 6
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Carbon-14
Atomic # 6
Mass # 14
 $\frac{+6}{+8} = 14$

Ion
-an atom or molecule where the number of **ELECTRONS** is different than the number of protons

Na⁺¹
before: Sodium 11p, 11e, 0n
after: Sodium 11p, 10e, 11n

Cl⁻¹
before: Chlorine 17p, 17e, 17n
after: Chlorine 17p, 18e, 17n

Student Handout- do the lesson quicker and easier

Atomic # +
Mass # N

Atomic # +
Mass # N